

Name \_\_\_\_\_

Date \_\_\_\_\_ Pd. \_\_\_\_\_

Chapter 6: Atomic Number and Mass Worksheet

**DIRECTIONS:** Using your Periodic Table, fill in the symbol, atomic number and atomic mass. Then figure out the amount of protons, neutron, and electrons. The elements with a negative sign(-, 2-, 3-) next to the symbol have extra electrons.

<u>Name</u>	<u>Symbol</u>	<u>Atomic Number</u>	<u>Atomic Mass</u>	<u># of Protons</u>	<u># of Neutrons</u>	<u># of Electrons</u>
Zinc						
	Hg					
		12				
Phosphorous						
		18				
	Cl					
Aluminum						
	Br					
		33				
Silver						
	S <sup>2-</sup>					
		28				

atomic mass = atomic weight

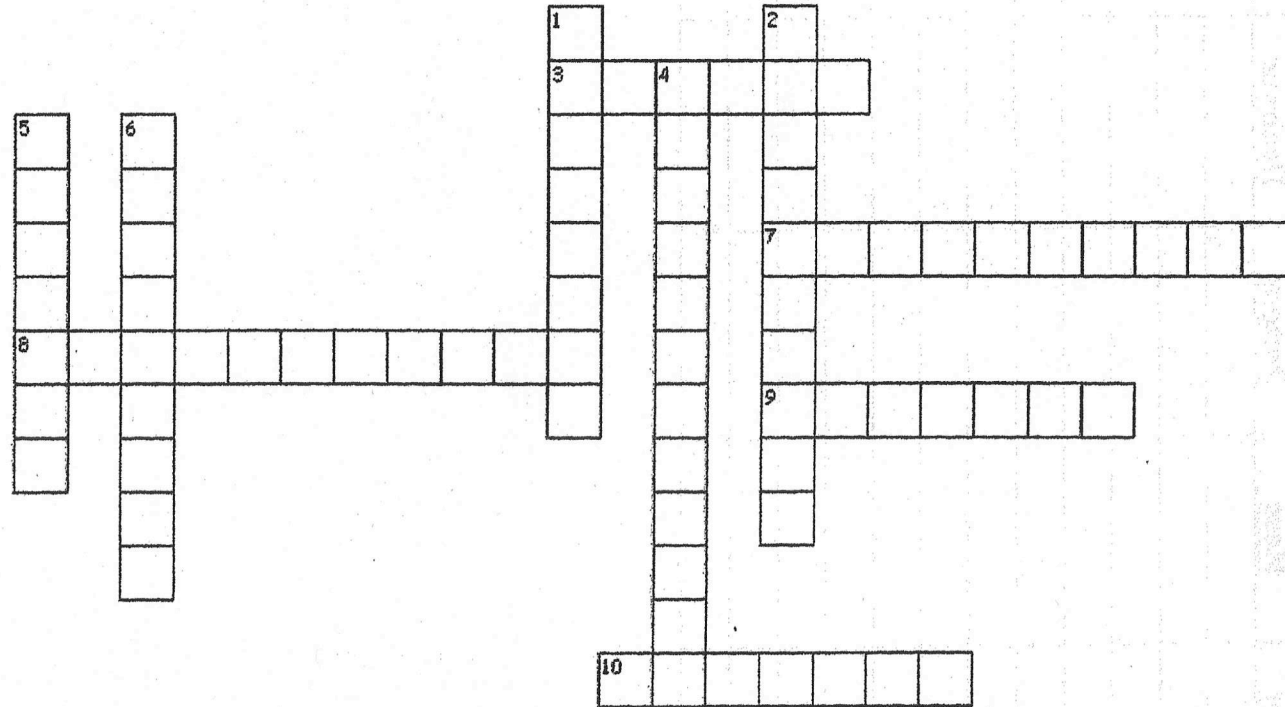
# of protons = atomic number

# of neutrons = atomic mass - atomic number

# of electrons = atomic number  
for neutral atoms

# Chapter 6 Crossword

Name \_\_\_\_\_



### Across

- 3. an atom with a perfect amount of protons and neutrons in the nucleus
- 7. the number of protons plus the number of neutrons in the nucleus of an atom
- 8. an atom's nucleus that has so many extra neutrons that the atom is unstable
- 9. when the atom or molecule has a greater amount of protons or electrons resulting in a positive or negative charge
- 10. the outside area around an atom's nucleus where electrons move

### Down

- 1. atoms with the same number of protons or atoms of the same element that have a different number of neutrons in the nucleus
- 2. the radioactive state when an atom's nucleus ejects an alpha particle
- 4. the number of protons in the nucleus of an atom
- 5. when 1 positive charge cancels out 1 negative charge resulting in a total charge of zero
- 6. the radioactive state when an atom splits a neutron into a proton and electron